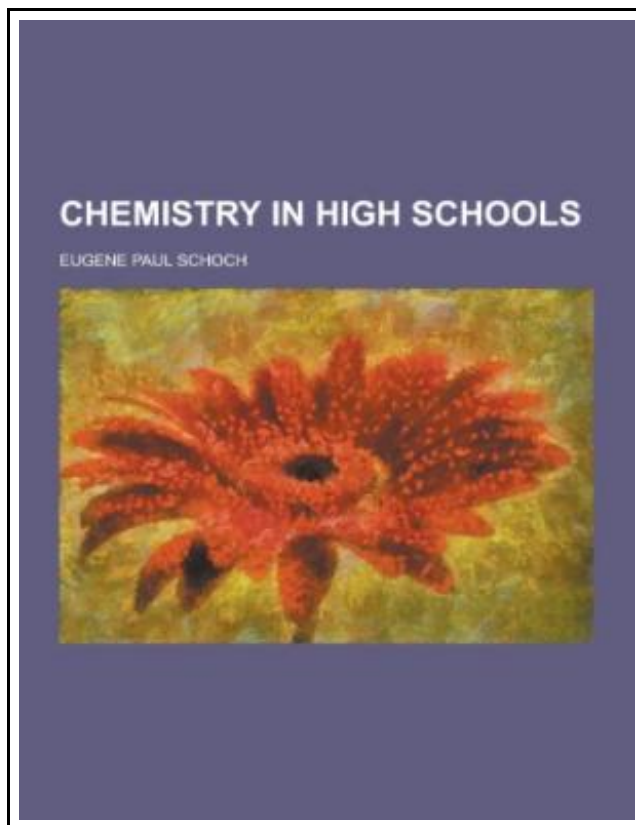


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Theclassics.Us, United States, 2013. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1911 edition. Excerpt: .and one other product (salt); while there is no action with the metals below hydrogen in the table. (b) Oxidizing acids (nitric acid, concentrated sulphuric acid) act on almost all metals except the lowest in the table (gold, platinum); but hydrogen does not appear as one of the products, and the action is altogether more complex than any action with non-oxidizing acids. The above general facts (a and b) should be learned and retained by the student. He should realize that by this means he learns a great number of important facts easily. In connection with (a) the following quantitative data is given: $\text{Fe} + 2\text{HCl} = \text{FeCl}_2 + 2\text{H}$ $\text{Zn} + 2\text{HCl} = \text{ZnCl}_2 + 2\text{H}$ $\text{Al} + 3\text{HCl} = \text{AlCl}_3 + 3\text{H}$ $\text{Fe} + \text{H}_2\text{SO}_4 = \text{FeSO}_4 + 2\text{H}$ $\text{Zn} + \text{H}_2\text{SO}_4 = \text{ZnSO}_4 + 2\text{H}$ $2\text{Al} + 3\text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + 6\text{H}$. These equations should be committed to memory and drilled on by means of numerical problems. 2. Preparation and collection of the gas to show some of its common properties. Perform as given in any text. The experiment shows that the gas is not very soluble in water, that it is lighter than air, that it forms water on burning and that when mixed with air in certain proportions it forms explosive mixtures. 3. Qualitative reduction of some metal oxide (lead oxide or copper oxide), Smith, Art. 20. Point out how this experimental procedure may be used (as it was used, by Dumas) to get the proportion by weight in which hydrogen and oxygen combine to form water. Incidentally the relative rate of diffusion of gases may be demonstrated, as usual, by means of...

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